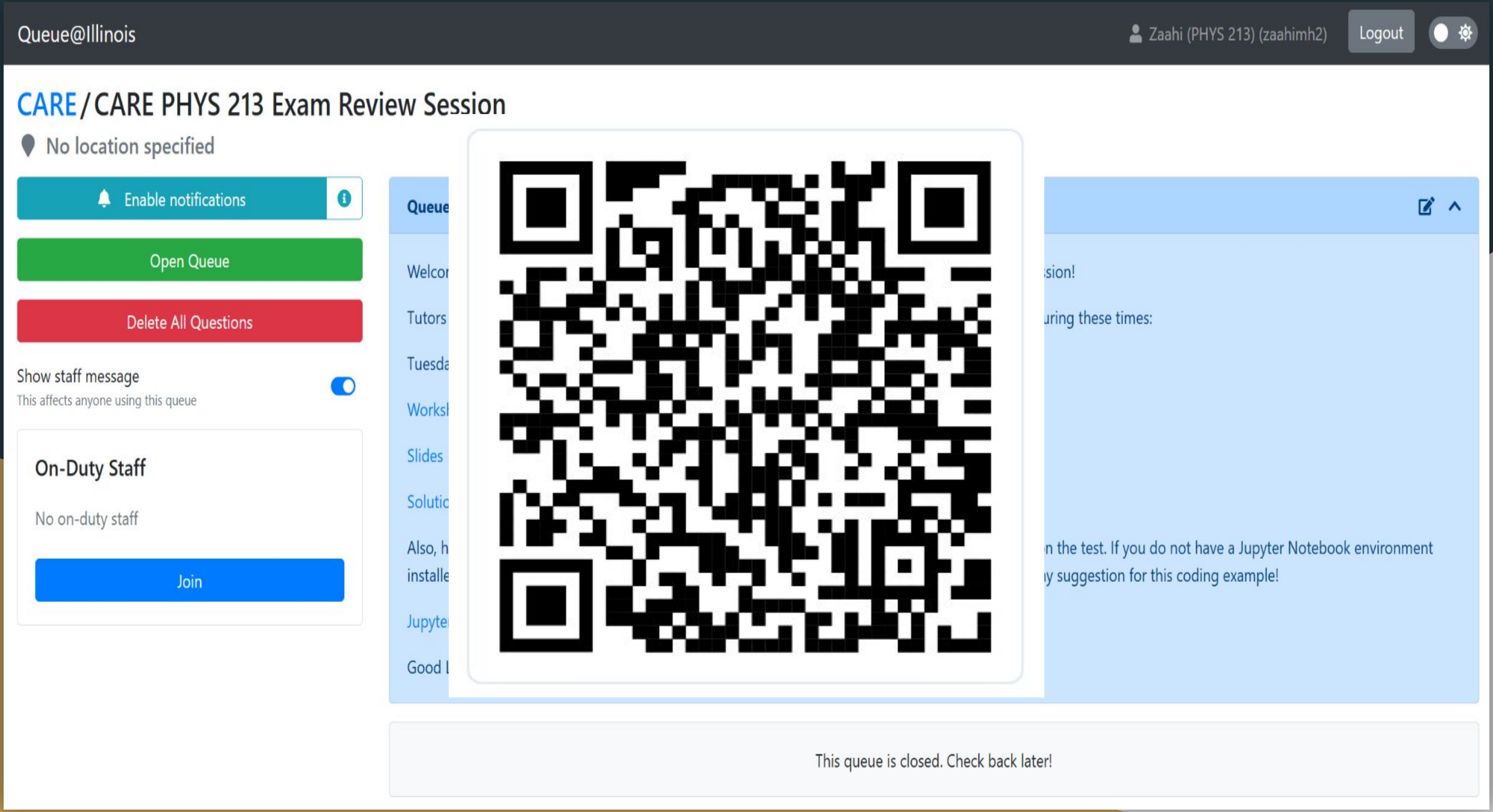
C.A.R.E. PHYS 213 Quiz 2 Review Session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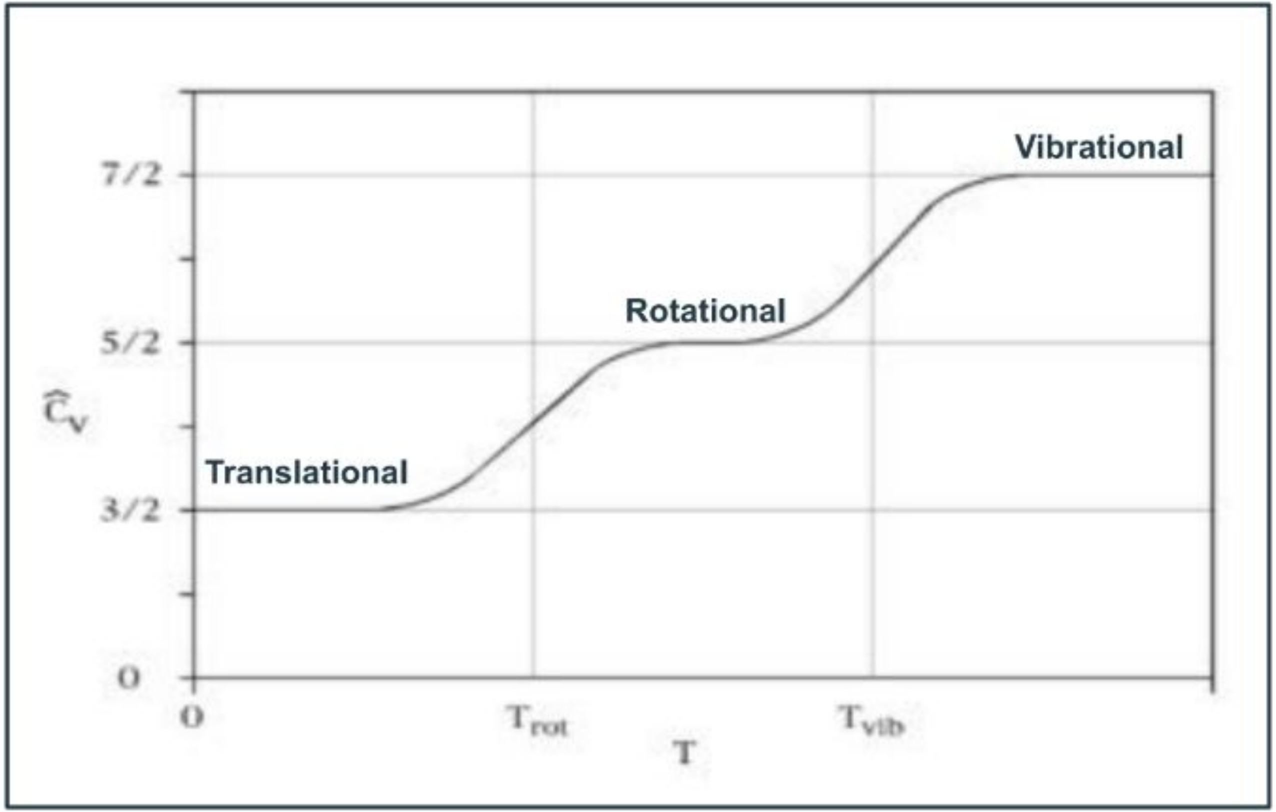
Units for the Exam

- Kinetic Theory of Ideal Gases
- Quasistatic Processes
- Thermodynamic Cycles
- Gibbs Free Energy

Ideal gas and Equipartition

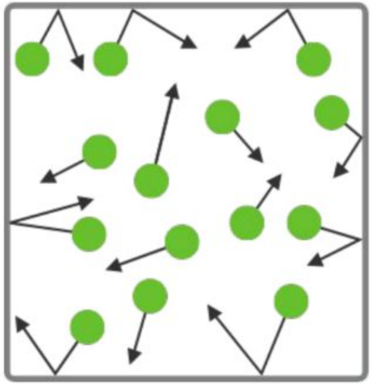
- Ideal Gas: Approximation of particles as points with no interactions:
  - Follows **ideal gas law**: $\boxed{pV = NkT = nRT}$

- **Equipartition:** each degree of freedom contributes ½ kT of energy
  - Internal energy in each particle:
    - $U = (N_{DOF}/2)kT$
  - Molar heat capacity:
    - $c_M = (N_{DOF}/2)kN_A$



Vibrational
Rotational
Translational
7/2
5/2
3/2
0
$\hat{C}_V$
0
T_{rot}
T_{vib}
T

Root-Mean-Square Velocity



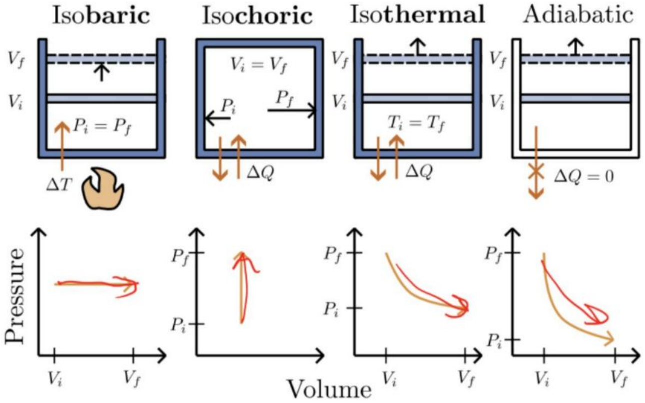
- v_{rms}: Average **(translational)** velocity of gas particles
- Translational Kinetic Energy: $KE_{translational} = 1/2\ m\ (v_{rms})^2$
- Relationship to temperature:

$$\boxed{\frac{1}{2}mv_{rms}{}^2 = \frac{3}{2}kT}$$

- This only applies to **ONE PARTICLE**
- **Notice:** this does **NOT** depend on the number of DOFs; it's **ALWAYS** $(3/2)kT$
  - Why? Translational KE only depends on the translational modes motion (there are only 3 translational modes: v_x, v_y, v_z)

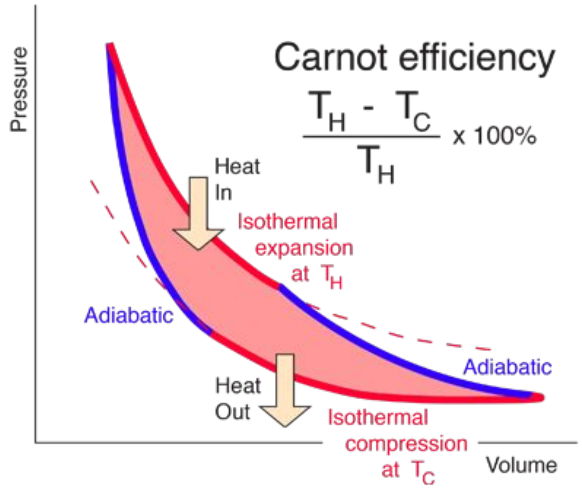
Thermodynamic Processes

- Isochoric or Isovolumetric
    - Constant VOLUME
- Isobaric
    - Constant PRESSURE
- Isothermal
    - Constant TEMPERATURE, REVERSIBLE $(\Delta S_{total} = 0, \Delta U = 0)$
- Adiabatic
    - Constant HEAT (dQ = 0), REVERSIBLE $(\Delta S_{total} = 0, \Delta Q = 0)$



Reversible Processes

- Isothermal + Adiabatic **processes** are **REVERSIBLE**
  - $\Delta S_{total} = 0$ (no change in entropy)
  - For **isothermal** processes:
    - PV = constant
  - For **adiabatic** processes:
    - PV^{γ} = constant
    - $\gamma = (2/N_{DOF}) + 1$ (given on equation sheet)



Example Problem - Adiabatic Process

- Assume we have a gas undergoing an adiabatic process, determine the work done given the following parameters:
  - $V_i = 10\ m^3$, $p_i = 10\ kPa$
  - $V_f = 4\ m^3$, $N_{DOF} = 3$

$$\gamma = \frac{2}{N_{DOF}} + 1 = \frac{2}{3} + 1 = \frac{5}{3}$$

$$pV^{\gamma} = Constant = p_i V_i^{\gamma} = (1000)(10)^{\frac{5}{3}} = 464158.88$$

$$W = \int_{V_i}^{V_f} p dV = \int_{V_i}^{V_f} \frac{C}{V_i^{\gamma}} dV = \left[C \frac{V_f^{\left(-\frac{5}{3}+1\right)}}{-\frac{5}{3}+1} \right] - \left[C \frac{V_i^{\left(-\frac{5}{3}+1\right)}}{-\frac{5}{3}+1} \right]$$

$$W = \left[(464158.88) \frac{4^{\left(-\frac{5}{3}+1\right)}}{-\frac{5}{3}+1} \right] - \left[(464158.88) \frac{10^{\left(-\frac{5}{3}+1\right)}}{-\frac{5}{3}+1} \right] \approx -126.30\ kJ$$

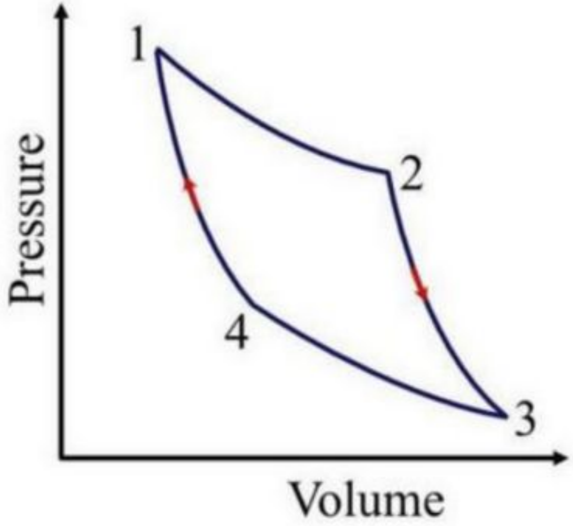
$$W \approx -126.30\ kJ$$

p-V Diagrams

$$W_{\text{by}} = \int_{V_i}^{V_f} p \, dV$$

- Used to visualize thermodynamic cycles

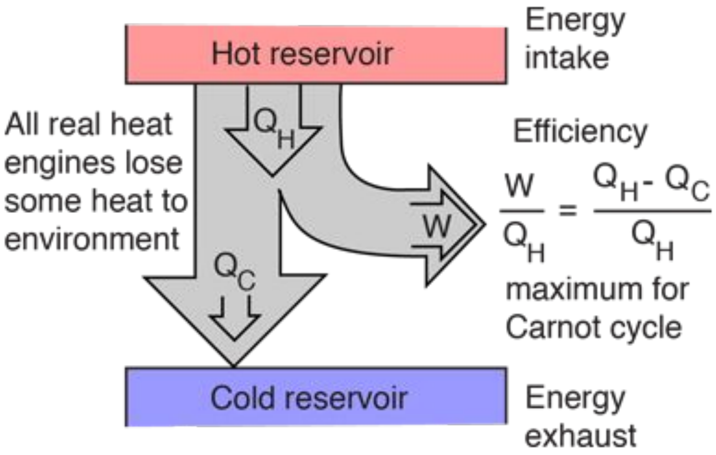
- **Area enclosed in the curve is equal to the work per cycle**
    - **Clockwise direction: work is positive** (engine did work)
    - **Counterclockwise direction: work is negative** (work done on engine)



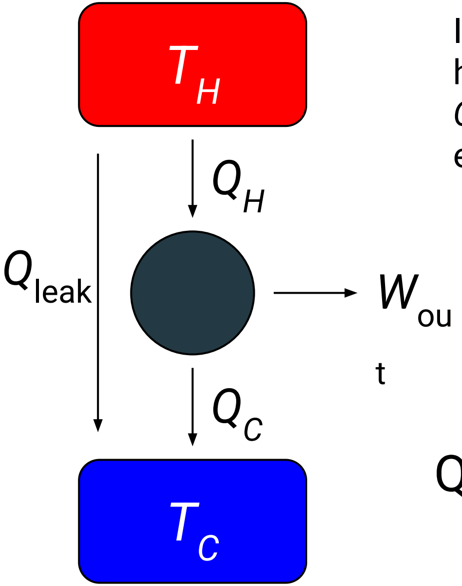
Heat Engines

Efficiency/COP can be thought of as "what you get out" divided by "what you put in."

- Cycles of Thermodynamic processes are used to make engines, heat pumps, and refrigerators
- **Efficiency of engines:** $\varepsilon = \dfrac{W_{by}}{Q_H} \leq 1 - \dfrac{T_C}{T_H}$
- **COP of pumps and refrigerators:**
  - Heat pump: $\mathrm{COP} = \dfrac{Q_H}{W_{on}} \leq \dfrac{1}{1 - \frac{T_C}{T_H}}$
  - Refrigerator: $\mathrm{COP} = \dfrac{Q_C}{W_{on}} \leq \dfrac{1}{\frac{T_H}{T_C} - 1}$



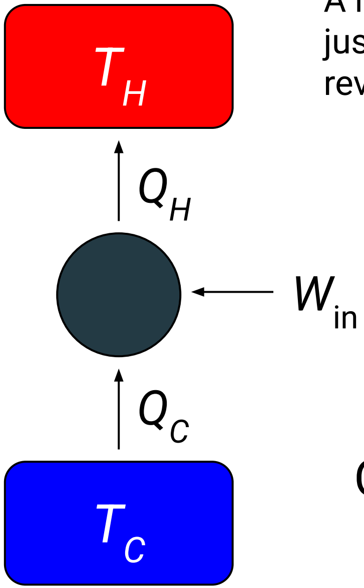
Engine, Pump, and Refrigerator Diagrams



Heat Engine Diagram

In a leaky engine, the heat we put in is Q_H + Q_{leak}, which affects the efficiency.

$$Q_H = W_{out} + Q_C$$



Heat Pump/Refrigerator Diagram

A heat pump or fridge is just a heat engine run in reverse.

$$Q_C = W_{in} + Q_H$$

Gibbs Free Energy

- Useful when **temperature** and **pressure** are fixed

$$\boxed{G = U - T_{env}S + pV}$$

- **Minimizing Gibbs of a system will maximize total (system + environment) Entropy**
  - As a system approaches equilibrium, free energy will decrease to a minimum

- **Fundamental Thermodynamic Relation in Equilibrium:**
  - $TdS = dU + pdV - \mu dN$
  - $\mu = (dG/dN) \rightarrow \mu N = G$ (at fixed temperature and pressure)
  - Equilibrium favors lowest μ

Good luck!

Feel free to ask any questions you may have.

You got this!