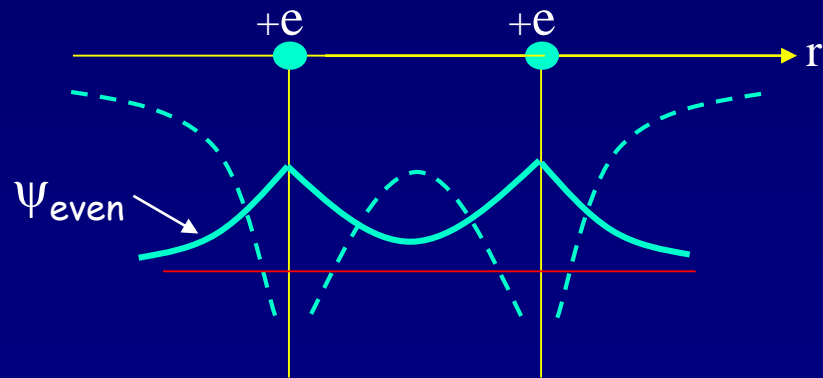
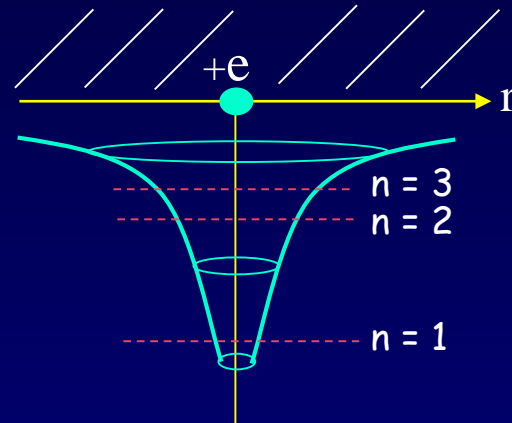


Lecture 19:

Building Atoms and Molecules



Today

Nuclear Magnetic Resonance

Using RF photons to drive transitions between nuclear spin orientations in a magnetic field

Far more useful than electron spin resonance (science, medicine)

Atomic Configurations

States in atoms with many electrons –
filled according to the Pauli exclusion principle → periodic table

Molecular Wave Functions: origins of covalent bonds

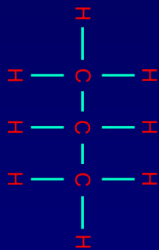
Chemistry: Example $\text{H} + \text{H} \rightarrow \text{H}_2$

Nuclear Magnetic Resonance

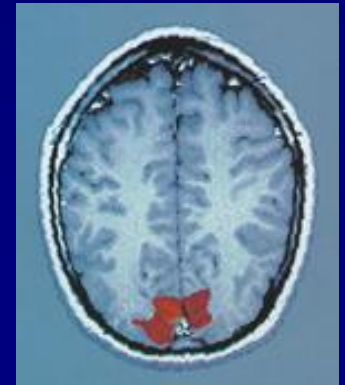
Not mass!

Just like electrons, the proton in the H atom also has a spin, which is described by an additional quantum number m_p , and therefore also a magnetic moment. However, it is several orders of magnitude smaller than that of the electron.

- The energy difference between the two proton spin states in a magnetic field is 660 times smaller than for electron spin states!
- But... There are many more unpaired proton spins than unpaired electron spins in ordinary matter. Our bodies have many unpaired protons in H_2O .
→ Signals large enough to be practically detected
- Also, energy difference depends a little bit on chemical (electronic) “environment,” so that spectroscopy → chemical analysis. Very useful!

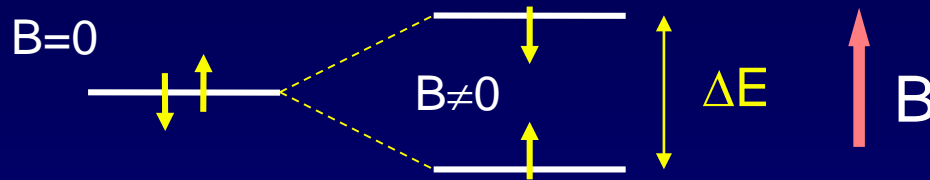


In order to image tissue of various types, **Magnetic Resonance Imaging** detects the small difference in the numbers of “up” and “down” **hydrogen proton spins** generated when the object studied is placed in a magnetic field. **Nobel Prize (2003): Lauterbur (UIUC)**



Example: Nuclear Spin and MRI

Magnetic resonance imaging (MRI) depends on the absorption of electromagnetic radiation by the nuclear spin of the hydrogen atoms in our bodies. The nucleus is a **proton with spin $\frac{1}{2}$** , so in a magnetic field **B** there are two energy states. The proton's magnetic moment is $\mu_p = 1.41 \times 10^{-26} \text{ J/Tesla}$.

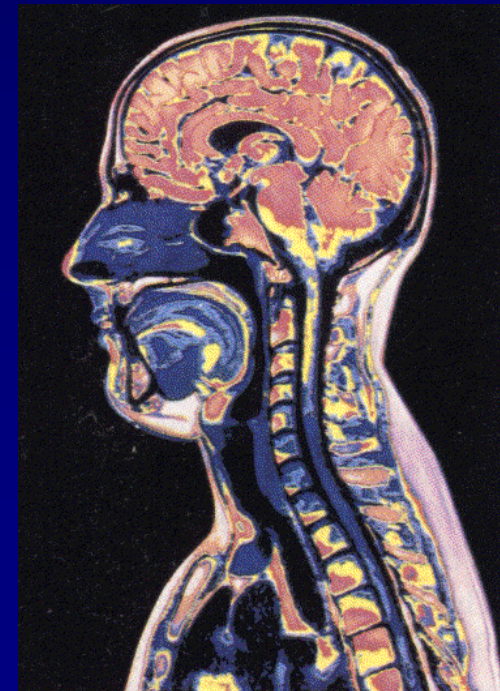


Chemical Shift: the electrons surrounding the protons change B a little bit, so ΔE depends a little on position within a molecule. Origin of CHEMICAL SHIFT

false color denotes proton chemical environment

1) The person to be scanned by an MRI machine is placed in a strong (1 Tesla) magnetic field. What is the energy difference between spin-up and spin-down proton states in this field?

2) What photon frequency, f , will be absorbed?



Act 1

We just saw that radio frequency photons can cause a nuclear spin to flip.
What is the angular momentum of each photon? ($S_{z,\text{proton}} = \pm \hbar/2$)

a. 0

b. $\hbar/2$

c. \hbar

Pauli Exclusion Principle

Let's start building more complicated atoms to study the Periodic Table.
For atoms with many electrons (e.g., carbon: 6, iron: 26, etc.) ...
What energies do the electrons have?

“Pauli Exclusion Principle” (1925)

No two electrons can be in the same quantum state. No two Fermions in same state

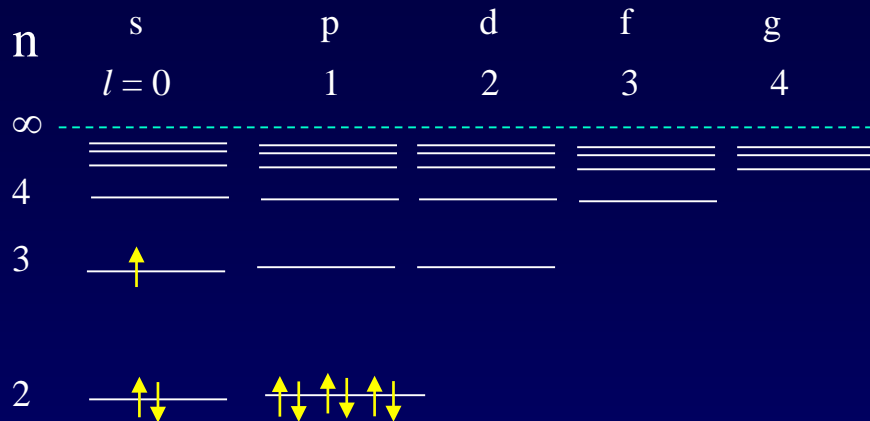
For example, in a given atom they cannot have the same set of quantum numbers n , l , m_l , m_s .

This means that each atomic orbital (n, l, m_l) can hold 2 electrons: $m_s = \pm 1/2$.

Important consequence:

- Electrons do not pile up in the lowest energy state.
It's more like filling a bucket with water.
- They are distributed among the energy levels starting with the lowest energy and sequentially adding electrons to higher energy states: obeying Exclusion Principle.
- Particles that obey this principle are called “fermions”.
Protons and neutrons are also fermions, but photons are not.

Filling Atomic Orbitals According to the Exclusion Principle

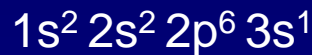


Energy

$$E_n = \frac{-13.6 \text{ eV}}{n^2} Z^2$$

In a multi-electron atom, the H-atom energy level diagram is distorted by Coulomb repulsion between electrons. Nevertheless, the H-atom diagram is useful (with some caveats) for figuring out the order in which orbitals are filled.

Example: Na ($Z = 11$)



l	label	#orbitals ($2l+1$)
0	s	1
1	p	3
2	d	5
3	f	7



Z = atomic number = # protons

Act 2

1. Which of the following states (n, l, m_l, m_s) is/are **NOT** allowed?

- a. $(2, 1, 1, -1/2)$
- b. $(4, 0, 0, 1/2)$
- c. $(3, 2, 3, -1/2)$
- d. $(5, 2, 2, 1/2)$
- e. $(4, 4, 2, -1/2)$

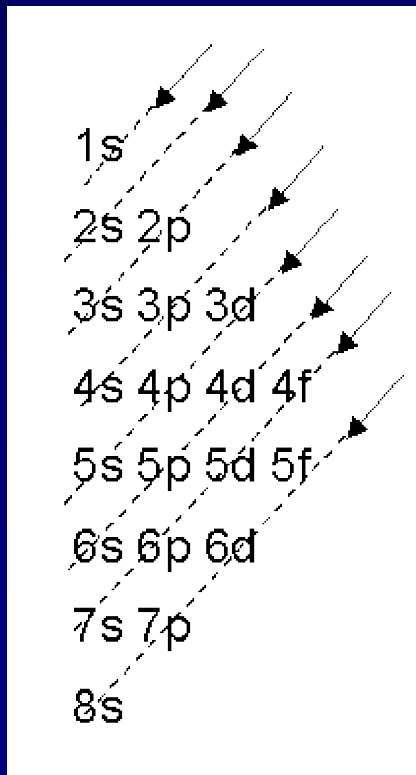
2. Which of the following atomic electron configurations violates the Pauli Exclusion Principle?

- a. $1s^2, 2s^2, 2p^6, 3s^2, 3d^8$
- b. $1s^2, 2s^2, 2p^6, 3s^2, 3d^4$
- c. $1s^2, 2s^2, 2p^8, 3s^2, 3d^8$
- d. $1s^1, 2s^2, 2p^6, 3s^2, 3d^9$
- e. $1s^2, 2s^2, 2p^3, 3s^2, 3d^{11}$

Filling Procedure for Atomic Orbitals

Due to electron-electron interactions, the hydrogen levels fail to give us the correct filling order as we go higher in the periodic table.

The actual filling order is given in the mnemonic table below. Electrons are added by proceeding along the arrows shown.



This is just a mnemonic.

Home exercise:

Bromine is an element with $Z = 35$. Find its electronic configuration (e.g., $1s^2 2s^2 2p^6 \dots$).

Note:

The chemical properties of an atom are determined by the electrons in the orbitals with the largest n , because they are on the “surface” of the atom.

Periodic Table of the Elements

	IA																															0
1	1 H	IIA																														2 He
2	3 Li	4 Be																														
3	11 Na	12 Mg	IIIB	IVB	VB	VIB	VII B	— VII —	IB	IB																						
4	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr														
5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe														
6	55 Cs	56 Ba	57 *La	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn														
7	87 Fr	88 Ra	89 +Ac	104 Rf	105 Ha	106 106	107 107	108 108	109 109	110 110																						

* Lanthanide Series

+ Actinide Series

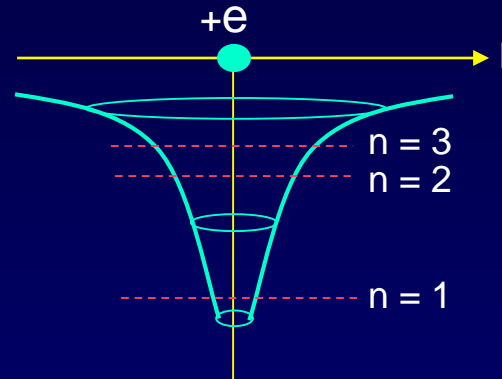
58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu
90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

As you learned in chemistry, the various behaviors of all the elements (and all the molecules made up from them) is all due to the way the electrons organize themselves, according to quantum mechanics.

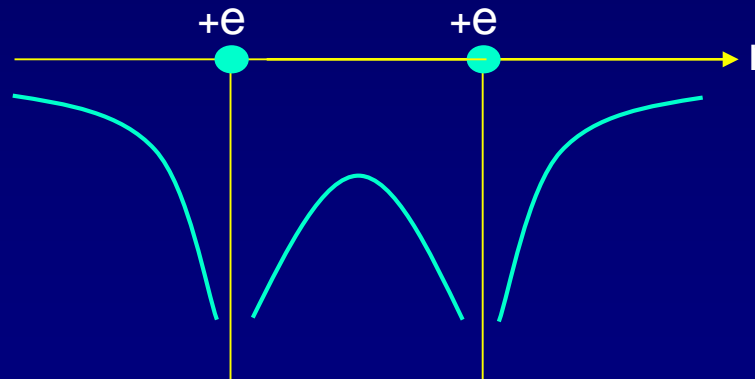
Bonding Between Atoms

How can two neutral objects stick together? $\text{H} + \text{H} \leftrightarrow \text{H}_2$

Let's represent the atom in space by its Coulomb potential centered on the proton (+e):



The potential energy due to the two protons in an H_2 molecule looks something like this:

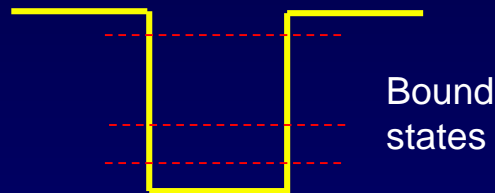


The energy levels for this potential are complicated, so we consider a simpler potential that we already know a lot about.

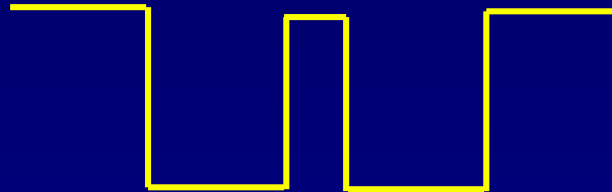
Particle in a Finite Square Well Potential

This has all of the qualitative features of molecular bonding, but is easier to analyze..

The 'atomic' potential:



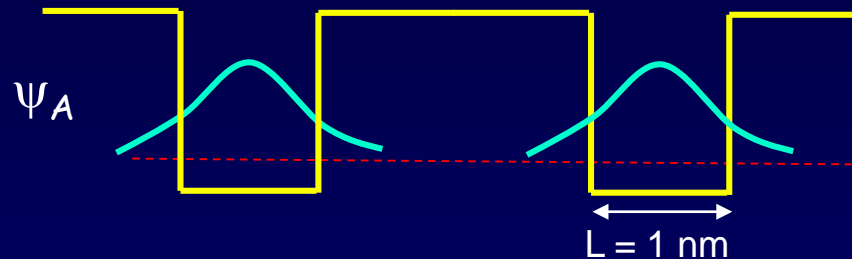
The 'molecular' potential:



Consider what happens when two “atoms” approach one other. There is one electron, which can be in either well (or both!). This is a model of the H₂⁺ molecule. We'll worry about the second electron later...

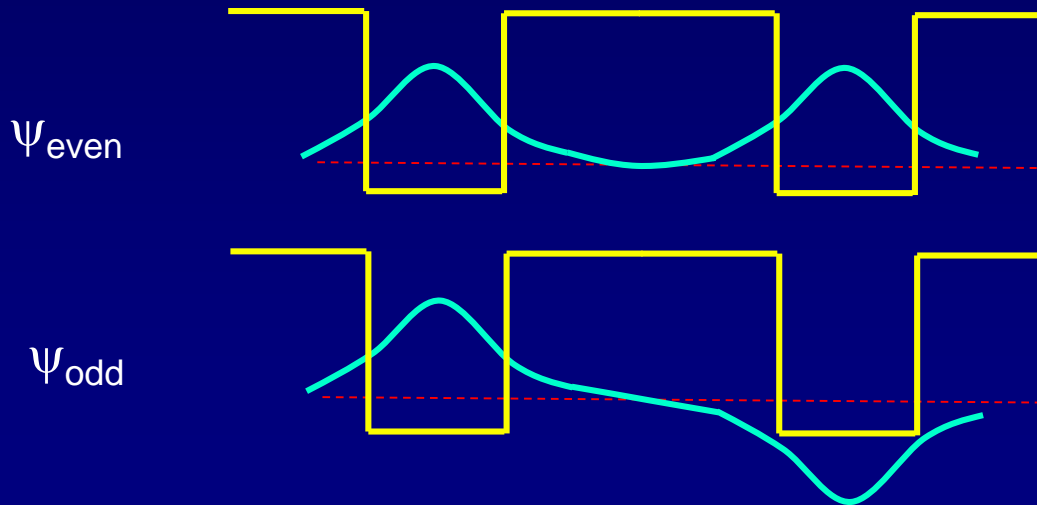
'Molecular' Wave functions and Energies

"Atomic" wave functions:



$$E \approx \frac{1.505 \text{ eV} \cdot \text{nm}^2}{(2L)^2} = 0.4 \text{ eV}$$

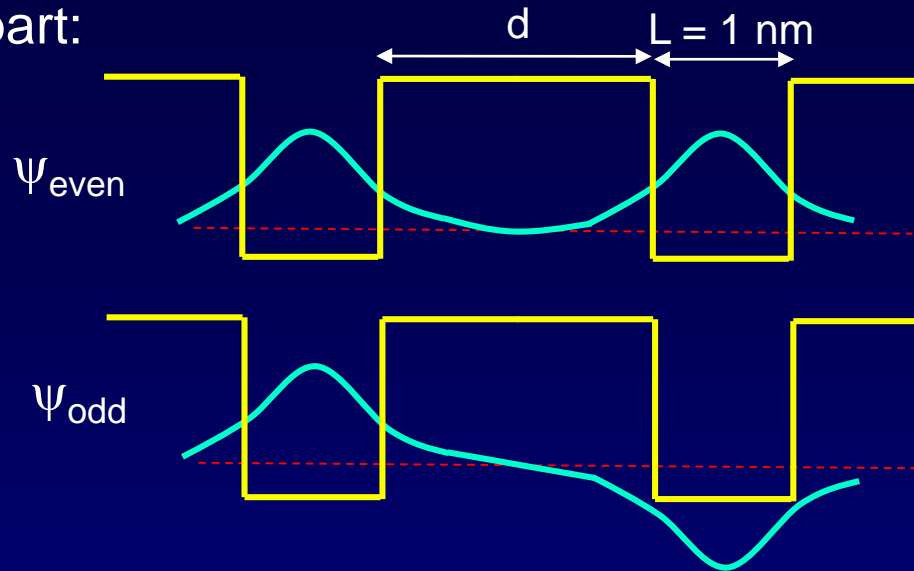
'Molecular' Wavefunctions: 2 'atomic' states \rightarrow 2 'molecular' states



When the wells are far apart, the 'atomic' functions don't overlap.
The single electron can be in either well with $E = 0.4 \text{ eV}$.

'Molecular' Wave Functions and Energies

Wells far apart:

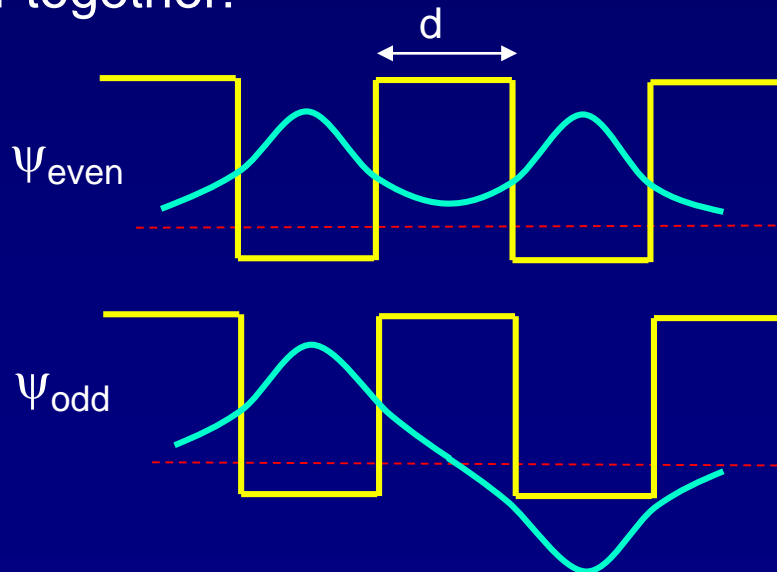


Degenerate states:
approx. infinite well energies

$$E \approx \frac{1.505 \text{ eV} \cdot \text{nm}^2}{(2L)^2} = 0.4 \text{ eV}$$

$$E \approx \frac{1.505 \text{ eV} \cdot \text{nm}^2}{(2L)^2} = 0.4 \text{ eV}$$

Wells closer together:

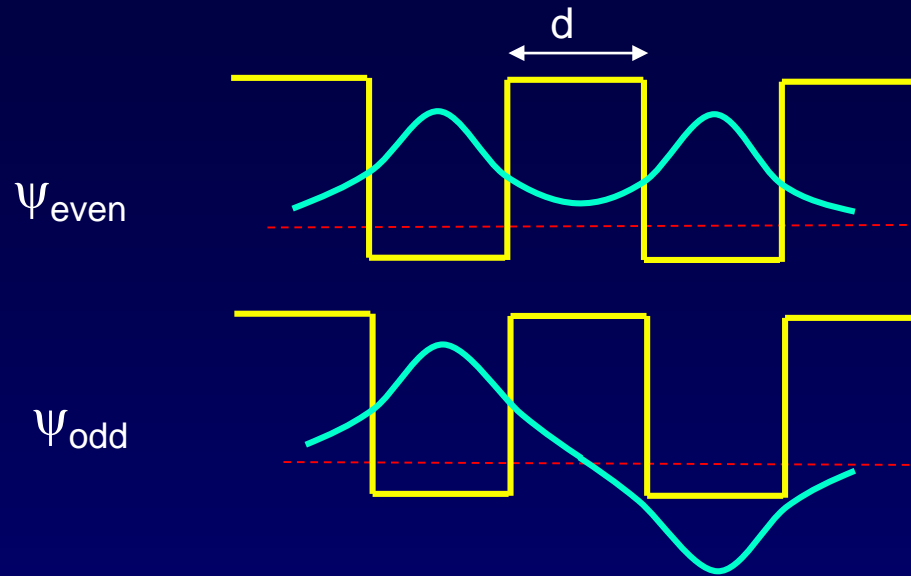


'Atomic' states are beginning to overlap and distort. ψ_{even} and ψ_{odd} are not the same. The degeneracy is broken:

$$E_{\text{even}} < E_{\text{odd}}$$

ψ_{even} : no nodes
 ψ_{odd} : one node

Act 3

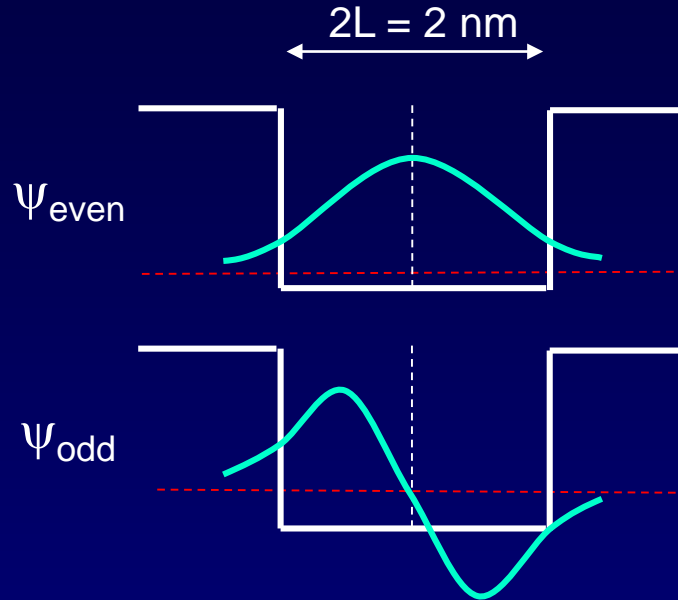


What will happen to the energy of ψ_{even} as the two wells come together (i.e., as d is reduced)? [Hint: think of the limit as $d \rightarrow 0$]

- a. E_{even} decreases.
- b. E_{even} stays the same.
- c. E_{even} increases.

Energy as a Function of Well Separation

When the wells just touch ($d = 0$, becoming one well) we know the energies:



approx. infinite well energies

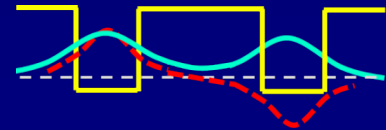
$$E_1 \approx \frac{1.505 \text{ eV} \cdot \text{nm}^2}{(4L)^2} = 0.1 \text{ eV}$$

($n = 1$ state)

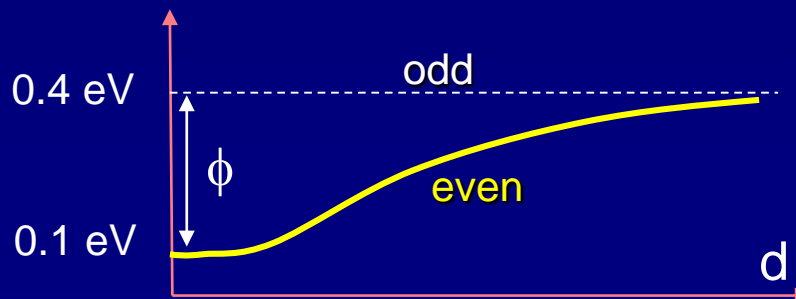
$$E_2 \approx \frac{1.505 \text{ eV} \cdot \text{nm}^2}{(4L)^2} \cdot 2^2 = 0.4 \text{ eV}$$

($n = 2$ state)

As the wells are brought together, the even state always has lower kinetic energy (smaller curvature, because it spreads out). The odd state stays at about the same energy. The node prevents it from spreading.



These numbers are not representative of real molecules. About 10 times too small.



Splitting between even and odd states:

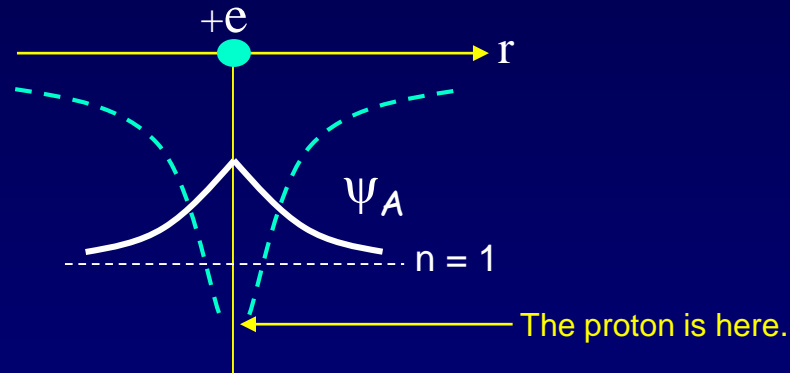
$$\Delta E = 0.4 - 0.1 \text{ eV} = 0.3 \text{ eV}$$

Molecular Wave Functions and Covalent Bond Energies with the Coulomb Potential

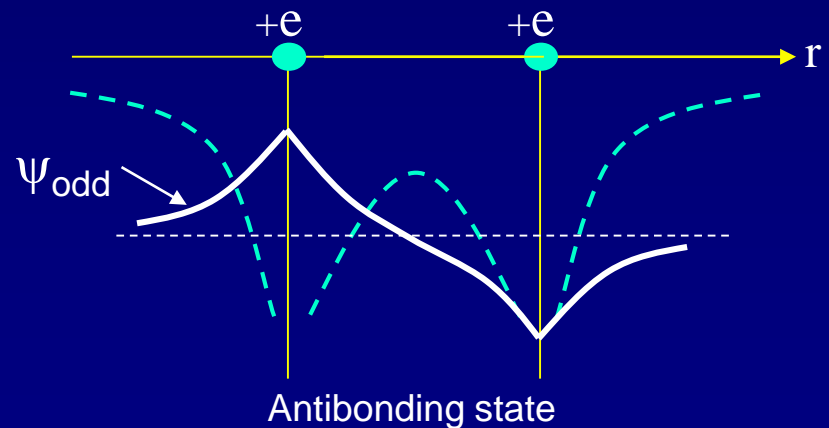
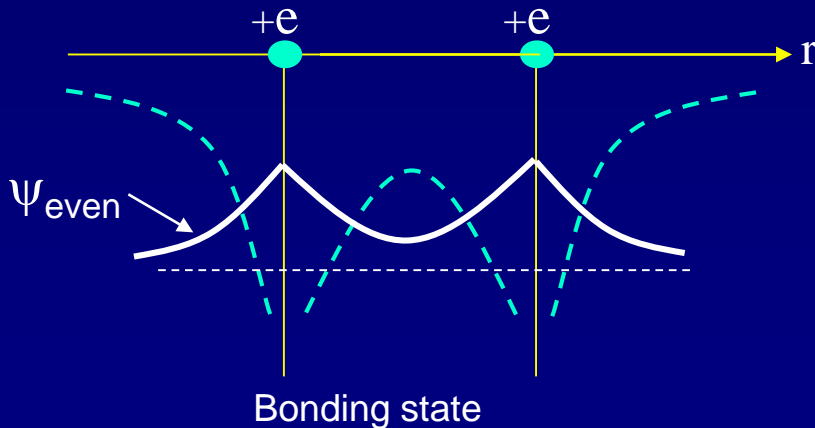
To understand real molecular bonding, we must deal with two issues:

- The atomic potential is not a square well.
- There is more than one electron in the well. Coulomb repulsion...

Atomic ground state (1s):



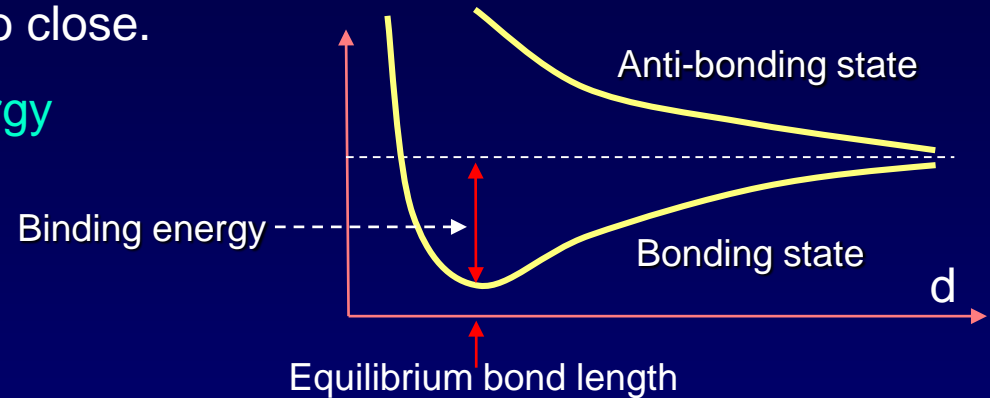
Molecular states:



Energy as a Function of Atom Separation

The even and odd states behave similarly to the square well, but there is also repulsion between the nuclei that prevents them from coming too close.

Schematic picture for the total energy of two nuclei and one electron:



Let's consider what happens when there is more than one electron:

- 2 electrons (two neutral H atoms): Both electrons occupy the bonding state (with different m_s). This is neutral H_2 .
- 4 electrons (two neutral He atoms). Two electrons are in the bonding state, and two electrons then must be added to the anti-bonding state \rightarrow ~zero energy savings coming from forming a covalent bond. Then repulsive force takes over, and the atoms don't stick. The He_2 molecule does not exist!

Summary

Nuclear Magnetic Resonance (magnetic moment due to nuclear spin)

Atomic configurations

- States in atoms with many electrons
- Filled according to the Pauli exclusion principle and Coulomb repulsion

Molecular wave functions: origins of covalent bonds

- Example: $\text{H} + \text{H} \rightarrow \text{H}_2$

Next time

Some practical uses of QM:

- Why do some solids conduct – others do not
- Solid-state semiconductor devices
- Lasers
- Superconductivity